

Homework #2

Due Wednesday, 12 September

Problem 1. *[stability and plume dispersion]*

Draw an atmospheric temperature profile that would result in the stack plume dispersions shown below.

a)



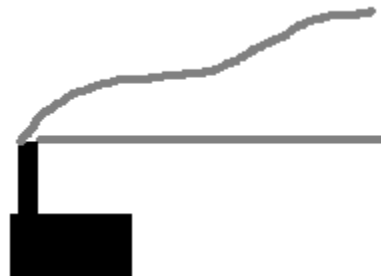
b)



c)



d)



e)



Problem 2. [atmospheric density / height relationships]

- (a) If a chemical species has a constant mixing ratio with height, what will its vertical profile in units of molecules cm^{-3} (or other concentration units) look like? Comment.
- (b) If a species has a constant mixing ratio profile with height, what does that imply about its sources and sinks?
- (c) If a species has a vertical mixing ratio profile that decreases with height, what does that imply about sources/ sinks? What if the profile increased with height?

Problem 3. [characteristic times]

- (a) Using the data below, compute characteristic times for several processes that affect tropospheric species "A".

Reaction rate constant with $\text{OH}\cdot = 1.5 \times 10^{-13} \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$

Dry deposition velocity (apply to 1 km boundary layer) = 0.6 cm s^{-1}

- (b) Compare these with characteristic times for exchange (see figure in your notes) and comment on the likely distribution of A in the atmosphere (trop. and strat.).

Problem 4. [Simple species budget]

Assume that the mass of methane in the atmosphere is represented by the variable M , and that methane has only one source, Q , (emissions of natural gas) and one sink, S (reaction with $\text{OH}\cdot$).

- (a) Sketch a simple representation of the budget of methane.
- (b) Derive a differential equation that describes the time rate of change of M . (Note that S depends upon M .)
- (c) Solve the differential equation for $M(t)$.
- (d) At steady state ($t \rightarrow \infty$), what is the value of M ?
- (e) Using typical values of $[\text{OH}\cdot]$ and the reaction rate constant for methane with $\text{OH}\cdot$ in the troposphere, $k = 6.3 \times 10^{-15} \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$, estimate the chemical lifetime of methane in the atmosphere. Given this lifetime, how would you expect methane concentrations to be distributed in the atmosphere (troposphere and stratosphere)?